

## Periodic Trends WS #1

#	Question	Answer
1	Where are the most active metals located?	
2	Where are the most active non-metals located?	
3	As you go from left to right across a period, does the atomic size decrease or increase. Why?	
4	As you travel down a group, the atomic size, does decreases or increases. Why?	
5	Is a negative ion is larger or smaller than its parent atom?	
6	Is a positive ion is larger or smaller than its parent atom?	
7	As you go from left to right across a period, does the first ionization energy generally decrease or increase? Why?	
8	As you go down a group, does the first ionization energy generally decrease of increase? Why?	
9	Where is the highest electronegativity found?	
10	Where is the lowest electronegativity found?	
11	Elements of Group 1A are called	
12	Elements of Group 2A are called	
13	Elements in the middle of the periodic table are called	
14	Group 7A elements are called	
15	Group 8A elements are called	
16	From left to right across the periodic table, do the elements go from (metals to nonmetals) or (nonmetals to metals)?	
17	The most active element in Group 7A is	
18	What orbitals are filling across the Transition Elements?	
19	Elements within a group have the same number of what?	
20	Are the majority of elements in the periodic table metals or non metals	
21	Elements in the periodic table are arranged according to their what?	
22	For each of the following sets of atoms, rank the atoms from smallest to largest atomic radius.	a) Li, C, F      b) Li, Na, K      c) Ge, P, O      d) C, N, Al      e) Al, Cl, Cu
23	For each of the following sets of atoms, rank them from lowest to highest ionization energy.	a) Mg, Si, S      b) Mg, Ca, Ba      c) F, Cl, Br      d) Ba, Cu, Ne      e) Si, P, He
24	For each of the following sets of atoms, rank them from lowest to highest electronegativity.	a) Li, C, N      b) Ne, C, O      c) Si, P, O      d) Mg, K, P      e) S, F, He