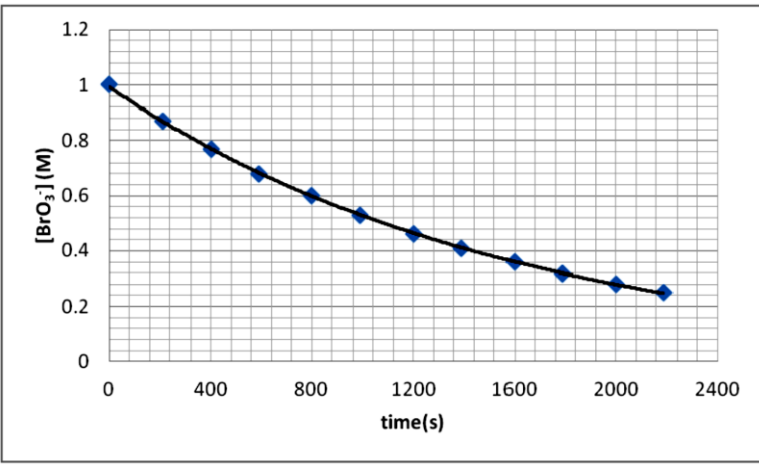


	Question	Steps																																
Sentence Format	<p>Ammonia (NH₃) is produced by the reaction between nitrogen and hydrogen gases. The concentration of ammonia increases from 0.257 M to 0.815 M in 15.0 min. Calculate the average rate of reaction over this time interval.</p>																																	
Chart Format	<p>Consider the reaction: $P_4 + 6 H_2 \rightarrow 4 PH_3$. A kinetics experiment was conducted and the following information was obtained. Calculate the rate of disappearance of H₂ between times 20 and 40 seconds.</p> <table border="1" style="margin-left: 20px;"> <thead> <tr> <th>Time (sec)</th> <th>[P₄]</th> </tr> </thead> <tbody> <tr> <td>0</td> <td>0.1 M</td> </tr> <tr> <td>20</td> <td>0.08 M</td> </tr> <tr> <td>40</td> <td>0.065 M</td> </tr> <tr> <td>60</td> <td>0.055 M</td> </tr> </tbody> </table>	Time (sec)	[P ₄]	0	0.1 M	20	0.08 M	40	0.065 M	60	0.055 M																							
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Graph Format	<p>The decomposition of BrO₃⁻ follows the following reaction: $BrO_3^- + 5Br^- + 6H^+ \rightarrow 3Br_2 + 3H_2O$ What is the average rate between time 400 sec and 2000 sec?</p>  <table border="1" style="margin-left: 20px;"> <caption>Data points from the graph</caption> <thead> <tr> <th>time(s)</th> <th>[BrO₃⁻] (M)</th> </tr> </thead> <tbody> <tr><td>0</td><td>1.0</td></tr> <tr><td>100</td><td>0.85</td></tr> <tr><td>200</td><td>0.75</td></tr> <tr><td>300</td><td>0.68</td></tr> <tr><td>400</td><td>0.62</td></tr> <tr><td>500</td><td>0.58</td></tr> <tr><td>600</td><td>0.55</td></tr> <tr><td>800</td><td>0.5</td></tr> <tr><td>1000</td><td>0.48</td></tr> <tr><td>1200</td><td>0.45</td></tr> <tr><td>1400</td><td>0.42</td></tr> <tr><td>1600</td><td>0.4</td></tr> <tr><td>1800</td><td>0.38</td></tr> <tr><td>2000</td><td>0.35</td></tr> <tr><td>2200</td><td>0.32</td></tr> </tbody> </table>	time(s)	[BrO ₃ ⁻] (M)	0	1.0	100	0.85	200	0.75	300	0.68	400	0.62	500	0.58	600	0.55	800	0.5	1000	0.48	1200	0.45	1400	0.42	1600	0.4	1800	0.38	2000	0.35	2200	0.32	
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