

Molarity Practice

- 1) An aqueous solution of NaOH contains 24 g of NaOH dissolved in 69 g of water. Find the molarity.
- 2) How many moles of HCl are present in 550 mL of 0.01 M HCl acid?
- 3) What is the molarity of a solution that contains 15.0 g NaCl in 1.25 L of solution?
- 4) What is the molarity of a 800 mL solution with 2.5 moles of solute?
- 5) What is the molarity of a 3L solution containing 65.0 g of Na₂CO₃ dissolved in H₂O?
- 6) How many grams of Al(OH)₃ are in 800 ml of a 0.2 M solution?
- 7) How many liters of a 0.3 M solution can be made by using 78 grams of isopropanol (C₃H₈O)?
- 8) What is the molarity of calcium hydroxide in this solution if I make a solution by adding 83 grams of calcium hydroxide to 750 mL of water?
- 9) What is the molarity of methanol in a solution if I make a solution by adding water to 35 mL of methanol (CH₃OH) until the final volume of the solution is 275 mL ? (The density of methanol is 0.792 g/mL)
- 10) Write your own molarity calculation. You must start with the name of an ionic compound but do not give the formula! That way you will have to show how to write the formula with crossing over. Start with a certain number of grams of that compound and add it to a certain number of mL of water. Then show how to solve the problem.